

CBS-Entrance Screening Test 2018

CBS-EST 2018

Information Brochure

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Syllabus

Entrance Test for Admission to

CENTER FOR BASIC SCIENCES

Pt. Ravishankar Shukla University

Raipur, Chhattisgarh



Academic Session 2018-19

Center for Basic Sciences

Pt. Ravishankar Shukla University
Raipur, Chhattisgarh

Introduction to CBS-EST 2018

The Entrance Screening Test (abbreviated EST) is a compulsory test for admission to 5-year Integrated MSc program in Basic Sciences, namely - Biology, Chemistry, Mathematics and Physics at the Center for Basic Sciences, Pt. Ravishankar Shukla University, Raipur. CBS is established by Pt. Ravishankar Shukla University, Raipur with the support of Government of Chhattisgarh, in 2015.

This institute has started with the mandate to provide high quality teaching in basic sciences by a faculty of distinguished scientists embedded in a vibrant research environment and to create a national pool of scientists ready to take up research challenges in the frontiers of basic and applied sciences. The Integrated MSc program at this institute follows a semester-based course structure and continuous assessment within a flexible and innovative academic curriculum, exposing the students to research early in their program.

The CBS is an institute equipped with state-of-art teaching and research laboratories, modern computational facilities and excellent libraries. All the students shall be accommodated at in-campus hostels, for both girls and boys, and are provided with an environment conducive to science education and research.

All the candidates admitted to 5-year Integrated M.Sc. program at CBS are eligible for scholarship of Rs. 5,000 per month. A contingency of Rs. 20,000 per year may be awarded for carrying out summer projects to some meritorious students.

The details of the Integrated M.Sc. program, courses, research activities, facilities and faculty profile at CBS can be found on their respective websites (hyper-linked with www.prsu.ac.in and www.cbsraipur.ac.in)

Eligibility criteria for admission

Educational qualification: Candidates seeking admission to CBS, Pt. Ravishankar Shukla University, Raipur for the academic session 2018-19 should be from science stream (having any combination of Biology/ Chemistry/ Mathematics/ Physics) at class XII and must write the CBS-EST 2018 examination. Candidates who have passed class XII examination or equivalent from any recognized Board in India or appearing in 2018 can appear in CBS-EST 2018 examination. Admission will be offered strictly on the basis of Merit List of CBS-EST 2018. Candidates securing at least 60% marks in aggregate or equivalent

grade in class XII examination will be eligible for admission to CBS. For candidates belonging to scheduled castes (SC), and scheduled tribes (ST) categories, the minimum requirement of marks is 55% in aggregate. Where only letter grades are available, a certificate from the Board specifying equivalent percentage of marks is required. In the absence of such a certificate, the decision of the Admissions Committee of the concerned Institution will be final.

Age limit: General category candidates born on or after July 01, 1998 are eligible for admission to the integrated M.Sc. program of CBS. The age limit is relaxed by 2 years for SC/ST/OBC candidates.

Necessary certificates supporting eligibility criteria have to be furnished at the time of admission. Please also note that the offer of admission is subject to verification of all original certificates at the time of counseling.

Number of seats & reservation

For the academic session 2018-19, the total number of seats at CBS is **60 (20 (PCB)+20(PCM) throughEST-2018&10 (PCB)+10(PCM) throughNEST-2018)**. Domicile Certificate is needed to be furnished during counseling. The number of seats reserved for SC, ST and OBC (Non-Creamy-Layer) is according to norms of Government of Chhattisgarh. To claim seats under reserved category, relevant documents must be furnished at the time of admission.

CBS-EST 2018 examination

The CBS-EST 2018 will be conducted at UTD, Pt. Ravishankar Shukla University, Raipur on June 11, 2018 (Monday), 11:00 am – 2:00 pm. Based on the performance in CBS-EST 2018, a common Merit List of the candidates will be prepared and posted on CBS website (www.cbsraipur.ac.in) on June 25, 2018 (Monday). The successful candidates will be called for counseling and admission will be strictly according to the Merit List until all the seats are filled. Separate merit list for different category candidates will be made available before counseling.

Examination rules: Candidates must reach the test centre at least half an hour before the start of the examination. The examination is of 3 hrs duration and will start at 11:00 am. Candidates will not be allowed to enter the examination hall after 11:30 am. The earliest a candidate can leave the examination hall is 12:30 pm, unless it becomes necessary to leave before on medical grounds. The earliest one can carry the question paper out is 12.30 pm. Use of log tables and calculators in the examination hall is not allowed. Candidates must bring their own pen, pencils (HB), eraser and sharpener. Exchange/ sharing of these items with other candidates is strictly prohibited. Candidates MUST bring their Admit Card and their school photo Identity Card or any photo-ID issued by Government agencies to the examination hall. Any candidate found adopting unfair means will be expelled from the examination hall without warning. Mobile phone and other similar electronic gadgets are strictly not allowed inside examination hall.

- 1. Question type:** Question paper will be in two parts 'Part-A: Compulsory' and 'Part-B: Optional'. 'Part-A: Compulsory' shall carry questions on Physics and Chemistry subjects. 'Part-B: Optional'

shall carry questions on Mathematics as well as Biology [Botany and Zoology] subject, out of which the candidate has to attempt/opt/choose only one subject, either Mathematics or Biology. Question paper will have 150 multiple choice objective type questions, each with four options, 50 questions in Physics, serially numbered from 1-50, 50 questions in Chemistry, numbering from 51-100 and 50 questions in Mathematics/Biology, numbering from 101-150.

2. **Marking Scheme:** Each correctly answered question will earn one mark, with a maximum of 150 marks for the paper. Question with no response indicated will not be awarded any mark and there will be no negative marking for all questions numbering 1-150.
3. More than one answer indicated against a question will be deemed as incorrect answer. Candidate will be required to choose the correct answer and mark in the OMR answer sheet by darkening the corresponding circle/ bubble against the serial number of the question with black/ blue ball-point pen.
4. **Language of the question paper:** Question paper for the examination shall be bilingual, i.e., in **ENGLISH and HINDI**. In case of any discrepancy in English and Hindi versions of question(s), English version will be taken as the final version.

Syllabus: The syllabus for CBS-EST 2018 primarily follows the NCERT/ CBSE/CG Board science syllabus of class XI-XII. The detailed syllabus for CBS-EST examination is provided in Annexure-2 (page 8).

How to apply

To apply for CBS-EST 2018, candidates can fill-up application form using online mode. For online application, candidates should visit www.cbsraipur.ac.in on or after April 25, 2018. Candidates are strongly advised to read through the detailed online application procedure available on the website.

Application Fee: The application fee for the candidates of General/OBC category is Rs.850. The application fee for SC/ST categories candidates is Rs.450. For online applications, it can be paid by using credit card/debit card/net-banking using online payment portal. Please refer to the fee-chart below.

Category	Online fee	Additional charges	Remark
General & OBC candidates	Rs. 850	money transfer charge and Vendor's service charge	Non-refundable
SC and ST Candidates	Rs. 450	money transfer charge and Vendor's service charge	Non-refundable

Mode of payment:

A multimodal payment portal where payments can be made using credit/debit cards and internet banking facility is attached with Online application portal. This payment method is available for online applicants. A detailed step-by-step instruction for payment through credit/debit cards and internet banking is available on CBS-EST website (www.cbsraipur.ac.in or hyper-link to www.prsu.ac.in) and also inside the online application portal for CBS-EST 2018. Candidates need not send any document to CBS-EST office (paperless application). Please refer to the instruction sheet (the 'How to Apply' tab on www.cbsraipur.ac.in) for details on all payment methods.

Please refer to the instruction sheet (the 'How to Apply' tab on www.cbsraipur.ac.in) for details on all payment methods. Admit card: The admit card for CBS-EST 2018 will be available for downloading from

June 05, 2018. Admit cards will not be dispatched for online applicants. Such applicants must download their admit cards from CBS-EST website (after login). Safe-keeping of the admit card is strongly advised because the successful candidates have to produce the admit card in original during the counseling and admission.

Address for correspondence

Requests for any CBS-EST 2017 related queries by postal mail must be sent at

Coordinator/Centre Superintend

CBS-EST 2018

Or

Director

CENTER FOR BASIC SCIENCES

Pt. Ravishankar Shukla University

Raipur – 492010 Chhattisgarh

For any queries requiring quicker response, contact at CBS-EST helpline: 0771-2262216.

or email: cbsprsu@gmail.com

Important Dates

- Start of Online application for CBS-EST 2018: April 25, 2018 (Monday)
- Closing of Online application: May 31, 2018 (Monday)
- Downloading of Admit Card begins: June 05, 2018
- CBS-EST 2018 examination: **June 11, 2018 (Monday) 11:00am – 2:00pm**
- Display of Answers Key on website: June 12, 2018 at 5.00 pm (Tuesday)
- Last date to receive objections (Dava Apatti) : June 15, 2018 (Friday)
- Declaration of results in CBS website: June 25, 2018 (Monday)
- Counseling for admission – UR seats: June 27, 2018
- Counseling for admission – Reserved seats: June 29, 2018
- Beginning of academic session: July 02, 2018

Important things to remember

- Candidates must reach the examination venue at least half an hour (30 minutes) before the start of the examination.
- Candidates will not be allowed to enter the examination hall half an hour (30 minutes) after the start of the examination, i.e. 11:30 am.
- Candidates will be allowed to leave the examination hall only after one hour and thirty minutes from the start of the examination, i.e. 12:30 pm.
- Use of log tables and calculators in the examination hall are not allowed. Candidates must bring their own pen, pencil, eraser and sharpener. Exchange/sharing of these items with other candidates is strictly prohibited.
- Candidates MUST bring their Admit Card and Identity Card to the examination hall.
- Mobile phone and other electronic gadgets are not allowed inside examination hall.

Check-list for online applications

- ✓ Name and mailing address with PIN-code are entered.
- ✓ Passport size photograph is uploaded properly/pasted at correct place
- ✓ Signature is uploaded properly/signed at proper place

Note. If any of the above is found missing, the application will be Considered incomplete and will be rejected.

Objection (Dava Apatti) Process

- Ambiguity in question/ or answer key may be objected through e-mail with scanned copy of supporting evidence for justification
- Concerning e-mail for submission of objection: cbsprsu@gmail.com

Counseling Process

- Time of registration to participate in Counseling : 10:30 am -12:00 noon
- For participation in counseling, candidate should bring all original documents- Mark Sheet, TC, Migration Certificate, Cast Certificate (for reserved category) etc. along with multiple copies of these documents
- After having selection for admission in CBS, candidate has to pay the admission fee on the same date.
- Admission fee in CBS is as prescribed by the University.

SYLLABUS FOR CBS-EST 2018

CENTER FOR BASIC SCIENCES

Pt. Ravishankar Shukla University, Raipur

Biology

Cell Biology

Cell theory and cells as unit of life. Basic concepts of biomolecules – Proteins, Carbohydrates, Lipids, Nucleic acids. Tools and techniques of cell studies - use of microscope and calibration (microscopy), elements of microscope. Biomembranes - transport mechanism, cellular respiration. Cell organelles - structure and functions. Discovery and structure of DNA, processes of replication, transcription, genetic code and translation. Principles of the basic techniques in molecular biology. Enzymes- catalysis, kinetics, activation energy, competitive and non-competitive inhibition.

Genetics and Evolution

Fundamentals of genetics and evolution. Evidences and theories of organic evolution. Organization of the heredity material in chromosomes. Equational division. Reduction division. Mitosis and meiosis compared and contrasted. Significance of meiosis. Mendel's laws of inheritance. Discovery of linkage, sex-linked inheritance. Crossing-over, stage at which crossing-over occurs. Neurospora genetics. Mutation - discovery, types of Mutation and Mutations in diploids. Role of mutations in evolution. Elaboration of Mendel's laws of inheritance. Monohybrid or Dihybrid crosses. Human genetics - human chromosomes, sex-determination, sex-linked inheritance.

Ecology

Physical and biological factors influencing organisms. Food chains, pyramids of numbers and biomass. Biological equilibrium. Interspecific associations. Bio-diversity. Flora and fauna. Basics of microbial systems, Ecosystems.

Humans and Environment

Soil, rainfall and temperature with reference to natural resources. Our natural resources - their uses and abuses. Environmental pollution and preventive measures. Biodiversity and conservation.

Biotechnology

Principles of recombinant DNA technology. Applications of biotechnology.

Biology of Animal systems

Digestive System - Modes of nutrition. Different vitamin compounds and their deficiencies. Structure of alimentary canal and associated glands, digestive enzymes and gastrointestinal hormones. Absorption of products of digestion, peristalsis, balanced diet.

Respiratory System - Gaseous exchange in animals. Structure of respiratory organs, mechanism of breathing, gaseous transport, tissue respiration.

Circulatory System - Open and closed systems. Functions of blood and lymph. Microscopic structure of blood and blood vessels. Structures and working of heart. Distribution of arteries and veins. Circulation of blood coagulation. Blood groups.

Excretory System - Elimination of nitrogenous waste. Osmoconformers and osmoregulators. Structure and function of kidney tubules. Arrangement of excretory organs.

Nervous System - General account of brain, spinal cord and nerves. Reflex actions (simple and conditioned). Sense organs (eye and ear).

Reproductive System - Sexual and asexual reproduction. General arrangement and functions of reproductive organs.

Developmental Biology - Basic features of development in animals. Types of eggs, fertilization, cleavage, blastula. Stem cells- definition, types, uses, advantages and disadvantages, induced pluripotent stem cells. Different hormones and their roles.

Diversity of Animal Life — Principles of classification, binomial nomenclature. General classification of animal phyla up to classes (invertebrates) and upto sub-classes / order (vertebrates), General characters of fishes, amphibians, reptiles, birds and mammals.

Immunology - Basics of immune mechanisms and diseases- active and passive immunity, T and B cell responses, antigen presentation, principles of vaccination, monoclonal antibodies and their uses, immunology of AIDS.

Biology of Plant systems

Anatomy and Physiology of Plants - Meristems. Plant growth and development. Internal and external regulators of growth and development in plant. Plant reproduction. Internal structure of root, stem, secondary growth and leaves. Xylem and Phloem - their cell elements and functions. Internal structure of dicot and monocot leaves. Photosynthesis - history, importance, factors and mechanism, stomatal mechanism, transpiration and respiration. Comparative study of dicot and monocot anatomy. Absorption and cellwater relations, transport of water and minerals, turgor and turgor movements. Significance of life-cycles with special reference to alternation of generations as exemplified in Funaria, Selaginella and Pinus (no structural details). Plant hormones.

Systematics - Principles of classical and new systematics. Binomial nomenclature. Familiarity with taxa. Plant breeding and tissue culture.

Chemistry

Physical Chemistry

Measurements in chemistry: SI units for fundamental quantities, significant figures in calculations.

Mole concept: Avogadro number and mole concept, molar masses, mole fraction, molarity, molality, percent composition, stoichiometry. Equivalent weight and normality. Calculations based on mole concept and stoichiometry of different reactions. Oxidation-reduction reactions.

Gaseous and liquid states: Absolute scale of temperature. Gas laws, ideal gas equation, real gases and deviation from ideality, liquefaction of gases, van der Waals equation. Kinetic theory of gases; average, root mean square and most probable velocities and their relation with temperature. Law of partial pressures. Vapour pressure. Diffusion of gases.

Atomic structure and chemical bonding: Bohr model, spectrum of hydrogen atom, quantum numbers. Wave particle duality, de Broglie hypothesis. Uncertainty principle. Orbitals and quantum numbers; shapes and energy of s, p and d orbitals. Electronic configurations of elements (up to atomic number 36), filling of orbitals - Aufbau principle. Pauli's exclusion principle and Hund's rule. Hybridization involving s, p and d orbitals. Atomic orbital overlap and chemical bonds; ionic, covalent and coordinate bonds; bond parameters. Orbital energy diagrams for homo-nuclear diatomic species. Lewis structures. Hydrogen bond. Polarity in molecules, dipole moment (qualitative aspects). VSEPR theory and shapes of molecules. Valence Bond Theory. Molecular orbital theory of homo-nuclear diatomic molecules (qualitative idea).

Thermodynamics: Thermodynamic states. First law of thermodynamics. Internal energy, work and heat, pressure-volume work. Enthalpy and enthalpy change, Hess's law, heat of - reaction, fusion and vapourization. Second law of thermodynamics, entropy, free energy, criterion of spontaneity.

Chemical equilibrium: Laws of chemical Equilibrium, law of mass action. Equilibrium constant – factors affecting equilibrium constant and its applications. Le Chatelier's principle - effect of concentration, temperature and pressure. Significance of ΔG and ΔG_0 in chemical equilibrium. Relationship of K and ΔG . Ionic equilibrium. Acids and bases (Bronsted and Lewis concepts), salts. K_a , K_b , K_w , degree of dissociation, pH and their relationships. Solubility product, common ion effect. Hydrolysis of salts. Buffer solutions.

Electrochemistry: Redox reactions and electrode potential, Electrochemical cells, Galvanic cells and cell reactions. Standard electrode potential. Nernst equation and its relation to ΔG and K. Electrochemical series, emf of galvanic cells. Electrolysis and Faraday's laws of electrolysis. Electrolytic conductance, specific, equivalent and molar conductivity, Kohlrausch's law. Concentration cells. Batteries (primary and secondary), fuel cells, corrosion.

Chemical kinetics: Rates of chemical reactions. Order of reaction, rate constant. First order and pseudo first order reactions. Factors affecting rate of reaction – concentration, temperature (Arrhenius equation), catalyst.

Solid state: Classification of solids, amorphous and crystalline solids, crystalline state, crystal lattice and unit cells; seven crystal systems (cell parameters a, b, c, α , β , γ), close packed structure of solids (cubic), packing in fcc, bcc and hcp lattices. Packing efficiency, nearest neighbours, ionic radii. Simple ionic compounds, Imperfection in solids, point defects. Electrical and magnetic properties, band theory of metals.

Solutions: Solution of solid and gas in liquid. Concentration of solution. Ideal and nonideal solutions. Colligative properties. Vapour pressure of solution, Raoult's law. Molecular weight determination from

lowering of vapour pressure, elevation of boiling point and depression of freezing point. Abnormal molecular mass, vant Hoff factor. Osmosis – Osmotic pressure, reverse osmosis.

Surface chemistry:

(a) Adsorption – Physisorption and chemisorptions. Factors affecting adsorption of gases on solids. Adsorption isotherm. Catalysis – homogeneous and heterogeneous, Activity and selectivity. Enzyme catalysis.

(b) Colloidal state – Types, preparation and properties of colloids. Tyndall effect, 13 Brownian movement, electrophoresis, coagulation. Application of colloids. Micelles.

Inorganic Chemistry

Classification of elements and periodicity in properties: Modern periodic table, classification of elements, periodic trends in properties of elements – valence, oxidation state, atomic/ionic radius, ionization energy, electron gain energy, electronegativity, valency, chemical reactivity. Diagonal relationship. Anomalous behaviours of Li, Be, B, C.

Hydrogen: Isotopes, preparation, isolation, properties and uses. Hydrides – ionic, covalent and interstitial. Properties of water and heavy water. Hydrogen peroxide – Preparation, structure, reactions, uses. Hydrogen as fuel cell.

s- Block elements (Alkali and alkaline earth elements) – General characteristics and trends in properties.

(a) Group 1: Preparation, properties and reactions of alkali metals with emphasis on chemistry of Na and K and their compounds - oxides, peroxides, hydroxides, carbonates, bicarbonates, chlorides and sulphates. Uses.

(b) Group 2: Preparation, properties and reactions alkaline earth metals with emphasis on the chemistry of Mg and Ca and their compounds - oxides, peroxides, hydroxides, carbonates, bicarbonates, chlorides and sulphates. Uses.

p- Block elements: General characteristics and trends in properties.

(a) Group 13: Chemistry of Boron and its compounds - borax, boric acid and diborane.

(b) Group 14, 15 and 16: Chemistry of carbon, sulphur, nitrogen and phosphorus. Allotropy. Chemistry of oxides and oxyacids of these elements. Phosphines, phosphorus chlorides, ammonia, peroxide and ozone; silicones, silicon tetrachloride and silicates.

(c) Group 17: Chemistry of halogens, chemistry of chlorine in detail. Interhalogen compounds. HX and oxyacids of halogens.

(d) Group 18: Isolation, properties and reactions of inert gases with emphasis on chemistry of Xenon.

d-Block elements: (Mainly 3d elements) General characteristics and trends in properties. Variable oxidation states and their stabilities, colour (excluding the details of electronic transitions) and calculation of spin-only magnetic moment. Catalytic properties. Interstitial compounds, alloy formation. Preparation and properties of potassium dichromate and permanganate.

f- Block elements: (mainly lanthanides) General characteristics and trends in properties. Variable oxidation states. Lanthanide contraction and its consequences. Coordination compounds: Nomenclature of mononuclear coordination compounds. Isomerism. Hybridization and geometries of mononuclear coordination compounds. Magnetic properties. Werner's theory, VBT, CFT.

Metals and metallurgy: Occurrence of metals. General methods of extraction involving chemical principles – thermodynamic, electrochemical and redox principles. General operation stages involved in

metallurgical operation. Metallurgy of p-block element (emphasis on Al). Metallurgy of Fe-triad (more emphasis on Fe metallurgy). Metallurgy of coinage metals (Cu, Ag with more emphasis on Cu). Refining.

Organic Chemistry

Basic concepts: Representation of organic compounds. Hybridisations of carbon. Sigma and pi-bonds. Shapes of simple organic molecules. Inductive effect, electromeric effect, resonance effect, hyperconjugation. Keto-enol tautomerism. Determination of empirical and molecular formulae (only combustion method). Hydrogen bond - definition and effect on physical properties of alcohols and carboxylic acids. Acidity and basicity of 14 organic acids and bases. Methods of purification of compounds.

Reactive intermediates: Homolytic and heterolytic bond cleavages. Formation, structure and stability of - carbocation, carbanion and free radical.

Isomerism: Structural and stereoisomerism. Geometrical isomerism. Chirality. Enantiomers. Optical isomerism of compounds containing up to two asymmetric centres, (R, S and E, Z nomenclature excluded). Racemic mixture. Conformations of ethane and butane (Newman projections).

Nomenclature: IUPAC nomenclature of simple organic compounds (only hydrocarbons, mono-functional and bi-functional compounds), including benzene derivatives. . Alkanes: Preparation, properties and reactions. Idea of homologous series Combustion and halogenation of alkanes. Mechanism of photohalogenation. Wurtz reaction.

Alkenes and Alkynes: Preparation, properties and reactions of alkenes and alkynes. Isomerization. Acidity of alkynes. Acid catalysed hydration of alkenes and alkynes (excluding the stereochemistry), Reactions of alkenes with KMnO_4 , sulphuric acid. Reduction of alkenes and alkynes. Preparation of alkenes and alkynes by elimination reactions (excluding stereochemistry). Electrophilic addition reactions of alkenes with X_2 , HX , HOX and H_2O ($\text{X}=\text{halogen}$). Markovnikov rule. Peroxide effect. Polymerization of alkenes. Addition reactions of alkynes. Metal acetylides. Ozonolysis

Aromatic compounds: Aromaticity. Huckel theory of aromaticity. Structure of benzene. Isomerism in substituted benzenes. Electrophilic substitution reaction on benzene General mechanism. Orientating influence of substituents in electrophilic substitution reaction of monosubstituted benzenes. Electrophilic substitution reactions of benzene and substituted benzenes - halogenation, nitration, sulphonation, Friedel-Crafts alkylation and acylation (No mechanism).

Haloalkanes (Alkyl halides): Preparation from alkanes, alcohols, olefins. Grignard reagents and their reaction with aldehydes/ketones/esters/nitriles. Nucleophilic substitution reactions of alkyl halides with different nucleophilic species. $\text{S}_{\text{N}}1$ and $\text{S}_{\text{N}}2$ reactions with mechanism. Halogen exchange reaction. Polyhalogen compounds.

Haloarenes: Nucleophilic aromatic substitution in haloarenes and substituted haloarenes (excluding Benzyne mechanism and Cine substitution).

Alcohols: Preparation from – olefins, alkyl halides, carboxylic acids, aldehydes/ketones. Hydroboration reaction. Dehydration, oxidation to aldehydes and ketones. Reaction with sodium, phosphorus halides, ZnCl_2/HX , H_2SO_4 . Identification of p-, sec- and tertalcohols. Uses of methanol and ethanol. Phenols:

Preparation of phenol from halobenzene, cumene and benzene sulphonic acid. Acidity. Reactions of phenols - halogenation, nitration, sulphonation, with Zn. Reimer-Tieman reaction, Kolbe reaction.

Ethers: Preparation by Williamson's Synthesis, dehydration of alcohols. Reaction with H_2O , HX .

Aldehydes and Ketones - Preparation of aldehydes and ketones from – Alcohols, olefins, acid chlorides, arylalkanes, nitriles, esters, Friedel-Crafts reaction. Reactions with – Alcohols, HCN , $NaHSO_3$. Reactions- oxidation, reduction, oxime and hydrazone formation. Aldol condensation, Perkin reaction. Cannizzaro reaction. Haloform reaction. Tests to distinguish aldehydes and ketones.

Carboxylic acids - Acidity and structure-acidity relationship. Preparation of acids. Preparation of amides, acid chlorides, esters and anhydrides. ester hydrolysis. Reactions of acids with - thionyl chloride, P -halides, ammonia, alkalis, metals, halogens, reducing agents. Decarboxylation. Halogenation.

Amines - Basicity and structure-basicity relationship. Identification of p -, sec - and $tert$ amines. Preparation of amines from - nitro compounds, nitriles, amides, haloalkanes/aromatic compounds. Reaction with – Acids, alkylating agents, acylating agents, nitrous acid. Diazotization of aromatic primary amines - Reactions of aromatic diazonium salts - azo coupling reaction, Sandmeyer and related reactions. Carbylamine reaction of p -amines.

Carbohydrate: Classification of carbohydrates. mono- and di- saccharides (glucose and sucrose). Characteristic tests. Structure of glucose. Reactions of glucose- Oxidation, reduction, hydroxylamine, HI , acetic anhydride. Cyclic structure of glucose. Structures of - Sucrose, maltose, starch and cellulose . Glycoside formation and hydrolysis of sucrose.

Amino acids and proteins: α -amino acids. General structure of peptides and proteins. Peptide bond. Characteristic tests. Separation of amino acids using physical properties. Denaturation of proteins. Enzymes.

Polymers: Classification. Homo and co-polymers, Addition and condensation polymerizations. Polythene, nylons, polyesters, Bakelite, melamine-formaldehyde, rubber – natural and synthetic.

Mathematics

Algebra

Algebra of complex numbers, addition, multiplication, conjugation, polar representation, properties of modulus and principal argument, triangle inequality, cube roots of unity, geometric interpretations. Quadratic equations with real coefficients, relations between roots and coefficients, formation of quadratic equations with given roots, symmetric functions of roots.

Arithmetic, geometric and harmonic progressions, arithmetic, geometric and harmonic means, sums of finite arithmetic and geometric progressions, infinite geometric series, sums of squares and cubes of the first n natural numbers.

Logarithms and their properties.

Permutations and combinations, Binomial theorem for positive integral index, properties of binomial coefficients. Matrices as a rectangular array of real numbers, equality of matrices, addition, multiplication

by a scalar and product of matrices, transpose of a matrix, determinant of a square matrix of order up to three, inverse of a square matrix of order up to three, properties of these matrix operations, diagonal, symmetric and skewsymmetric matrices and their properties, solutions of simultaneous linear equations in two or three variables.

Addition and multiplication rules of probability, conditional probability, Bayes Theorem, independence of events, computation of probability of events using permutations and combinations.

Trigonometry

Trigonometric functions, their periodicity and graphs, addition and subtraction formulae, formulae involving multiple and sub-multiple angles, general solution of trigonometric equations.

Relations between sides and angles of a triangle, sine rule, cosine rule, half-angle formula and the area of a triangle, inverse trigonometric functions (principal value only).

Analytical geometry

Two dimensions - Cartesian coordinates, distance between two points, section formulae, shift of origin. Equation of a straight line in various forms, angle between two lines, distance of a point from a line. Lines through the point of intersection of two given lines, equation of the bisector of the angle between two lines, concurrency of lines. Centroid, orthocentre, incentre and circumcentre of a triangle.

Equation of a circle in various forms, equations of tangent, normal and chord. Parametric equations of a circle, intersection of a circle with a straight line or a circle, equation of a circle through the points of intersection of two circles and those of a circle and a straight line.

Equations of a parabola, ellipse and hyperbola in standard form, their foci, directrices and eccentricity, parametric equations, equations of tangent and normal. Locus Problems.

Three dimensions - Direction cosines and direction ratios, equation of a straight line in space, equation of a plane, distance of a point from a plane.

Differential calculus

Real valued functions of a real variable, into, onto and one-to-one functions, sum, difference, product and quotient of two functions, composite functions, absolute value, polynomial, rational, trigonometric, exponential and logarithmic functions.

Limit and continuity of a function, limit and continuity of the sum, difference, product and quotient of two functions, L'Hospital rule for evaluation of limits of functions.

Even and odd functions, inverse of a function, continuity of composite functions, intermediate value property of continuous functions. Derivative of a function, derivative of the sum, difference, product and quotient of two functions, chain rule, derivatives of polynomial, rational, trigonometric, inverse trigonometric, exponential and logarithmic functions.

Derivatives of implicit functions, derivatives up to order two, geometrical interpretation of the derivative, tangents and normals, increasing and decreasing functions, maximum and minimum values of a function, Rolle's Theorem and Lagrange's Mean Value Theorem.

Integral calculus

Integration as the inverse process of differentiation, indefinite integrals of standard functions, definite integrals and their properties, Fundamental Theorem of Integral Calculus.

Integration by parts, integration by the methods of substitution and partial fractions, application of definite integrals to the determination of areas involving simple curves.

Formation of ordinary differential equations, solution of homogeneous differential equations, separation of variables method, linear first order differential equations.

Vectors

Addition of vectors, scalar multiplication, dot and cross products, scalar triple products and their geometrical interpretations.

Physics

General: Units and dimensions, dimensional analysis. least count, significant figures. Methods of measurement (Direct, Indirect, Null) and measurement of length, time, mass, temperature, potential difference, current and resistance.

Design of some simple experiments, such as: i) Searle's method to determine Young's modulus, ii) determination of 'g' by simple pendulum, iii) speed of sound using resonance tube, iv) coefficient of friction using angle of repose, v) determination of focal length of a convex lens by plotting a graph between 'u' and 'v', vi) refractive index of material of prism using the method of minimum deviation, vii) verification of Ohm's law, viii) resistance of galvanometer using half deflection method, ix) specific heat of a liquid using calorimeter, x) I-V characteristic curve for p-n junction in forward and reverse bias.

Graphical representation and interpretation of data. Errors in the measurements and error analysis.

Mechanics: Kinematics in one and two dimensions (Cartesian coordinates only), projectiles. Uniform circular motion. Relative velocity. Newton's laws of motion. Inertial and uniformly accelerated (linear only) frames of reference. Static and dynamic friction. Kinetic and potential energy. Linear and circular simple harmonic motion. Work and power. Conservation of linear momentum and mechanical energy.

Systems of particles. Centre of mass and its motion. Centre of gravity. Impulse. Elastic and inelastic collisions.

Law of gravitation. Centripetal acceleration and centrifugal force. Gravitational potential and field. Acceleration due to gravity. Motion of planets and satellites in circular orbits. Escape velocity.

Rigid body, moment of inertia, parallel and perpendicular axes theorems, moment of inertia of uniform bodies with simple geometrical shapes. Angular momentum, Torque. Conservation of angular momentum. Dynamics of rigid bodies with fixed axis of rotation. Rolling without slipping of rings, cylinders and spheres. Equilibrium of rigid bodies. Collision of point masses with rigid bodies.

Hooke's law and stress – strain relations. Elastic limit, plastic deformation. Young's modulus, bulk and shear moduli.

Pressure in a fluid. Pascal's law. Buoyancy. Surface energy and surface tension, capillary rise. Viscosity – Stoke's and Poiseuille's law, Terminal velocity. Qualitative understanding of turbulence. Reynolds number. Streamline flow, equation of continuity. Bernoulli's theorem.

Sound and mechanical waves: Plane wave motion, longitudinal and transverse waves, superposition of waves. Progressive and stationary waves. Vibration of strings and air columns. Resonance (qualitative understanding). Beats. Speed of sound in gases. Doppler Effect.

Thermal physics: Thermal expansion of solids, liquids and gases. Calorimetry, latent heat. Heat conduction in one dimension. Elementary concepts of convection and radiation. Newton's law of cooling. Ideal gas laws. Specific heats (CV and CP for monoatomic and diatomic gases). Isothermal and adiabatic processes, bulk modulus of gases. Equivalence of heat and work. First and second law of thermodynamics and its applications (only for ideal gases). Entropy. Blackbody radiation - absorptive and emissive powers. Kirchhoff's law. Wien's displacement law, Stefan's law.

Electricity and magnetism: Coulomb's law. Electric field and potential. Electrical potential energy of a system of point charges and of electrical dipoles in a uniform electrostatic field; Electric field lines. Flux of electric field. Gauss's law and its application in simple cases, such as to find field due to infinitely long straight wire. Uniformly charged infinite plane sheet and uniformly charged thin spherical shell.

Capacitance - Calculation of capacitance with and without dielectrics. Capacitors in series and parallel. Energy stored in a capacitor.

Electric current. Ohm's law. Series and parallel arrangements of resistances and cells. Kirchhoff's laws and simple applications; Heating effect of current.

Biot-Savart's law and Ampere's law. Magnetic field near a current carrying straight wire, along the axis of a circular coil and inside a long straight solenoid. Force on a moving charge and on a current carrying wire in a uniform magnetic field.

Magnetic moment of a current loop. Effect of a uniform magnetic field on a current loop. Moving coil galvanometer, voltmeter, ammeter and their conversions.

Electromagnetic induction - Faraday's law, Lenz's law. Self and mutual inductance. RC, LR and LC circuits with and A.C. Sources.

Optics: Rectilinear propagation of light. Reflection and refraction at plane and spherical surfaces, Deviation and dispersion of light by a prism. Thin lenses. Combination of mirrors and thin lenses. Magnification. Wave nature of light - Huygen's principle, interference limited to Young's double slit experiment. Elementary idea of diffraction – Rayleigh criterion. Elementary idea of polarization – Brewster's law and the law of Malus.

Modern physics: Atomic nucleus. Alpha, beta and gamma radiations. Law of radioactive decay. Decay constant. Half-life and mean life. Binding energy and its calculation. Fission and fusion processes. Energy calculation in these processes.

Photoelectric effect. Bohr's theory of hydrogen like atoms. Characteristic and continuous X-rays, Moseley's law. de Broglie wavelength of matter waves. Heisenberg's uncertainty principle.